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## Reports of the Royal College of Chemistry; And Researches Conducted in the Laboratories in the Years 1845-6-7

By Royal College of Chemistry

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Excerpt: .This salt is only slightly soluble in cold water, and even less so in alcohol and ether. It contains equal equivalents of cyaniline and of the hydride of nitric acid. When burnt with oxide of copper, the following results were obtained: I. 0-3244 grm. of the salt gave: 0 5504 carbonic acid, and 0-1415 water. II. 0-5292 the salt gave: 0-9019, carbonic acid, and 0-2131 water, which numbers correspond to the per centage composition: Carbon 46-27 46-48 Hydrogen. 480 4-47 leading to the formula: Cy, C13 H7 N, H N06. as the following comparison with the theoretical numbers will show: 14 equivs. of Carbon. 8 - - Hydrogen. 3, Nitrogen. 6 - - Oxygen. 1 equiv. of Nitrate of Cyaniline Nitrate of cyaniline forms a fine double salt with nitrate of silver. Platinum compound of cyaniline.--A dilute solution of the hydrochlorate of cyaniline gives no...



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